Figure 2.7 A Cathode-Ray Tube

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Figure 2.8 Deflection of Cathode Rays by an Applied Electric Field

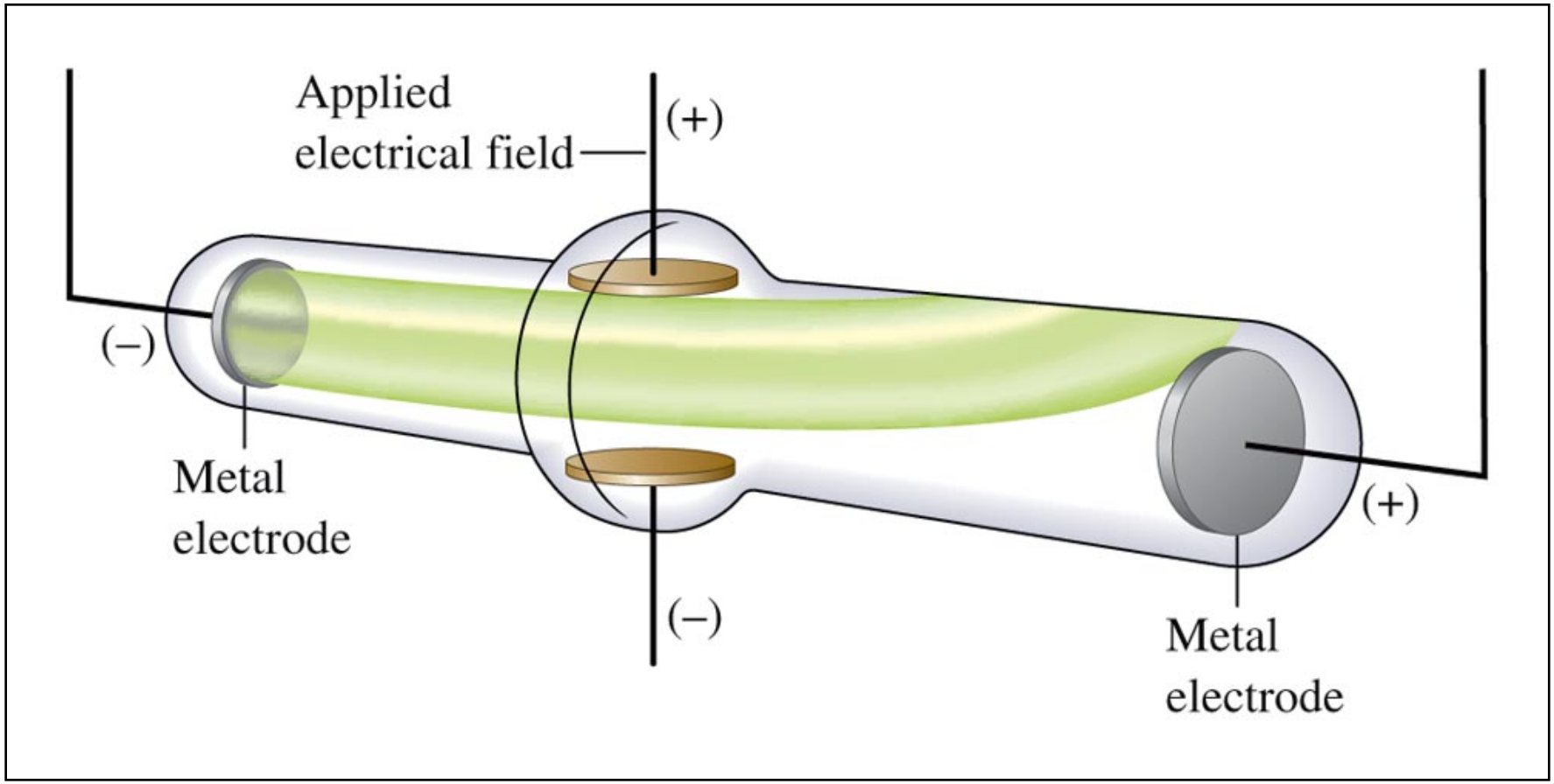

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Figure 2.9 The Plum Pudding Model of the Atom

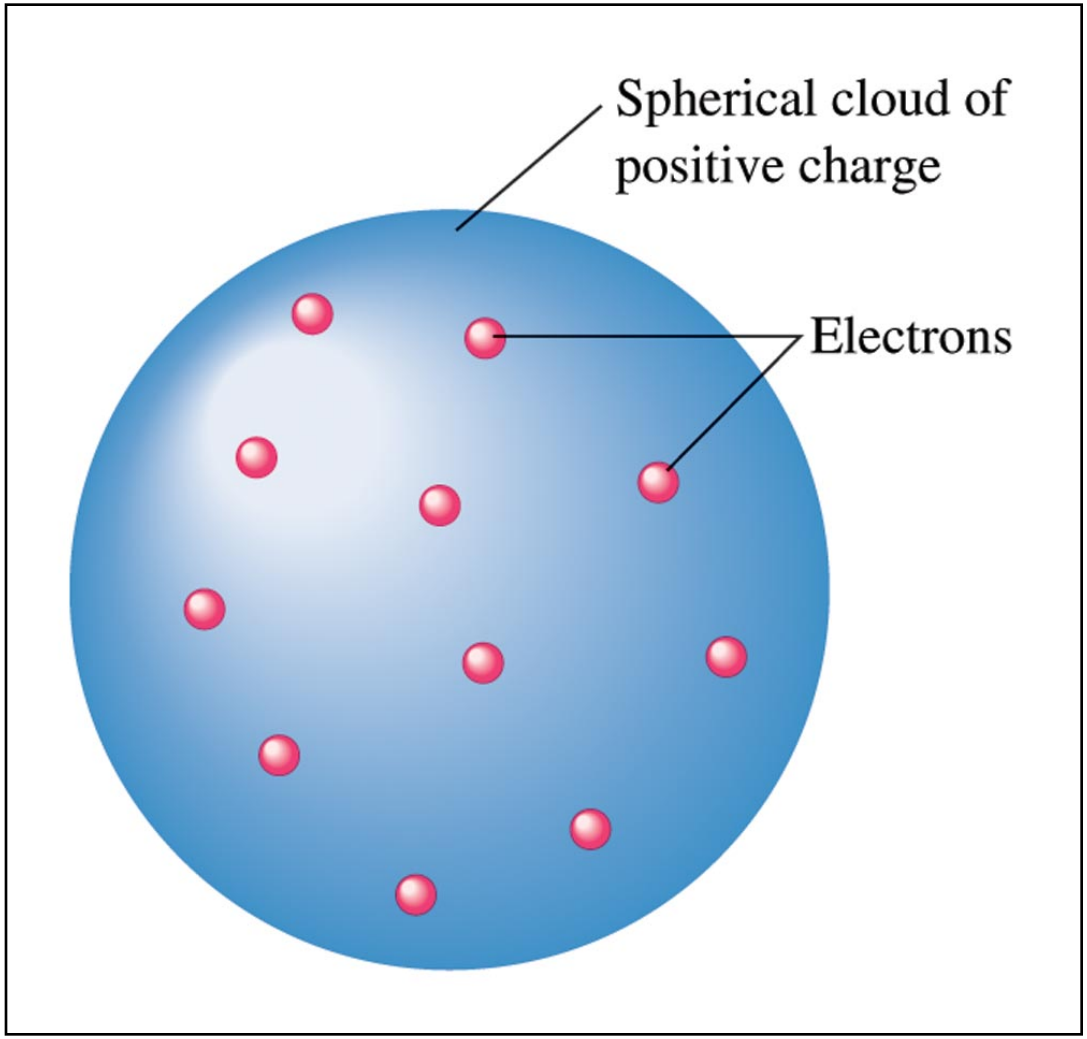

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Figure 2.10 A Schematic Representation of the Apparatus Millikan Used to Determine the Charge on the Electron

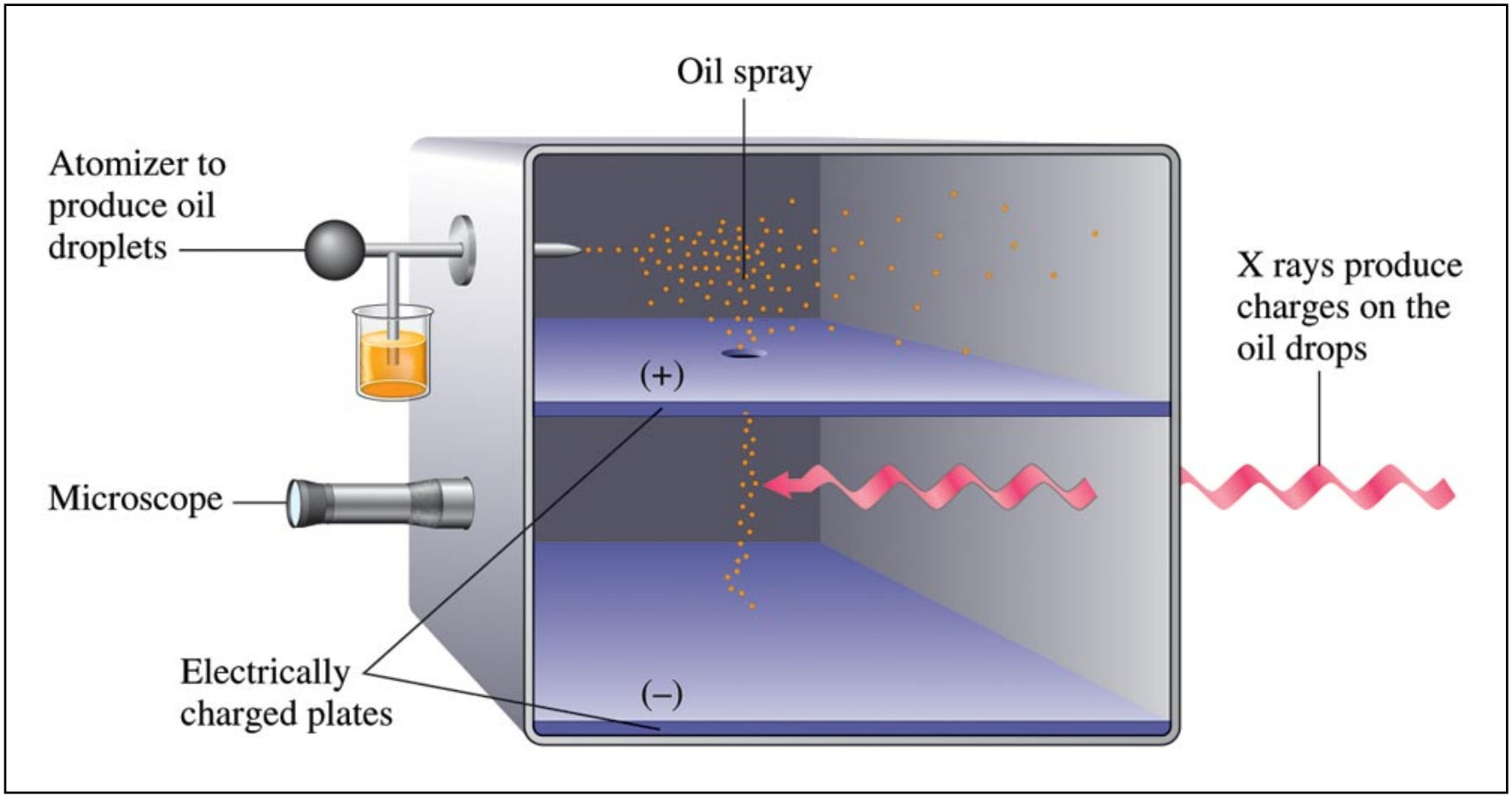

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Figure 2.12 Rutherford's Experiment On *a*-Particle Bombardment of Metal Foil

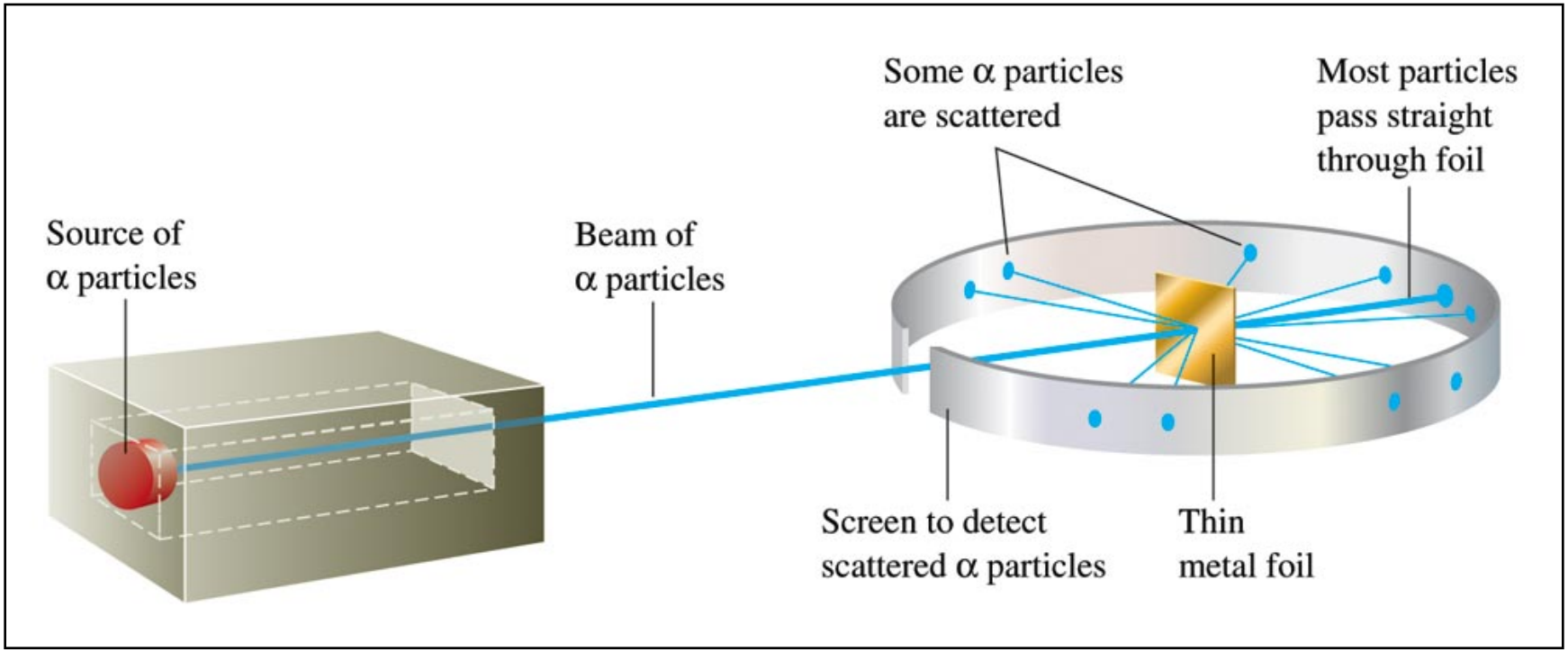

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Figure 2.13 a & b
(a) Expected Results of the Metal Foil Experiment if Thomson's Model Were Correct
(b) Actual Results

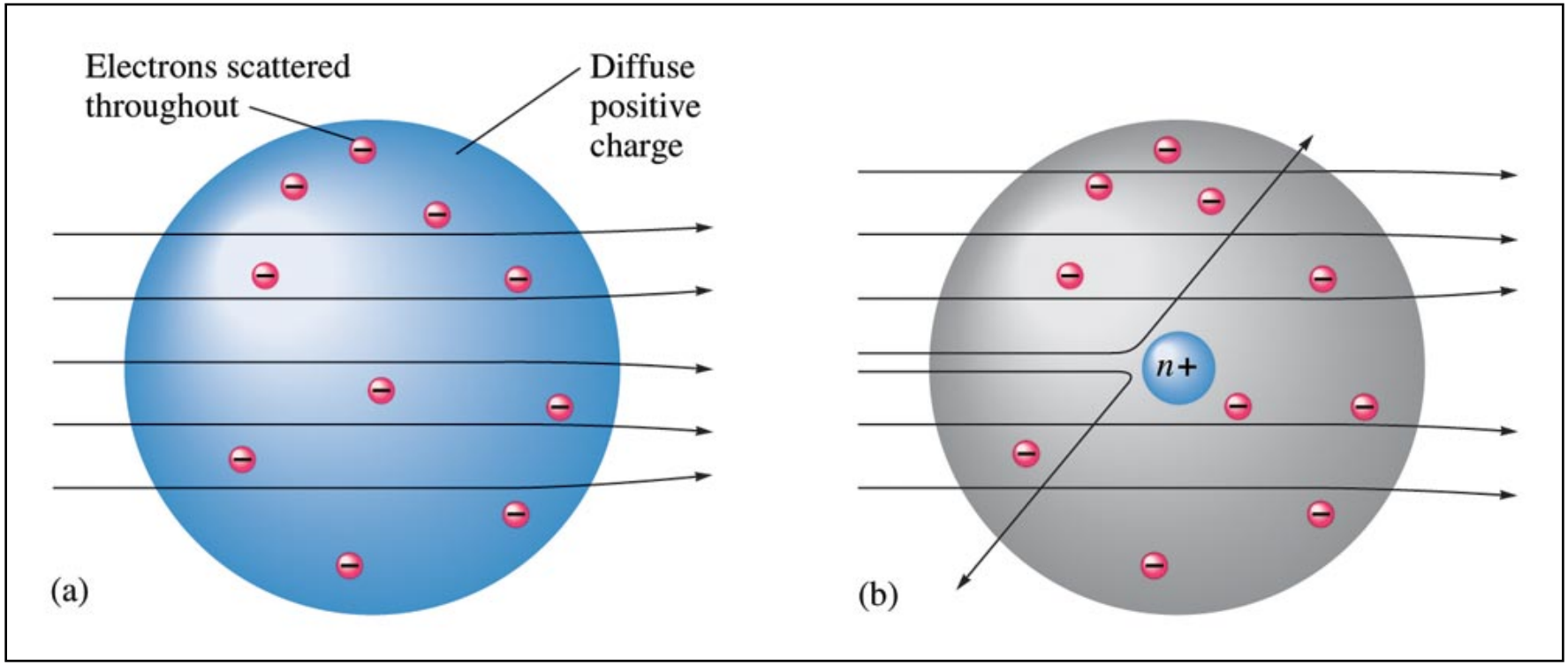

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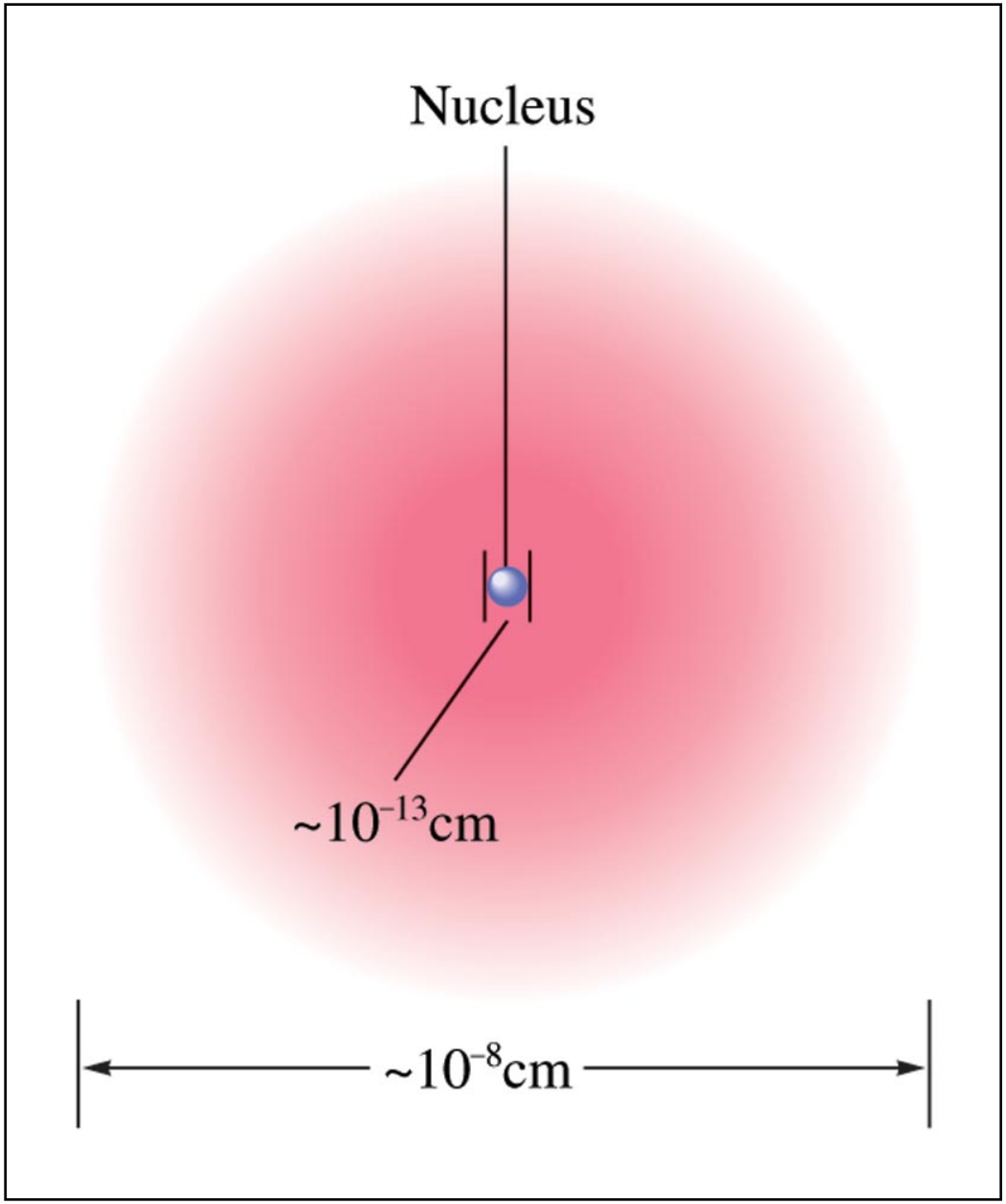
Figure 2.14 Cross Section of a Nuclear Atom



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Atomic Nucleus

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Table 2.1 The Mass and Charge of the Electron, Proton, and Neutron

TABLE 2.1 The Mass and Charge of the Electron, Proton, and Neutron

Particle	Mass	Charge*
Electron	9.11×10^{-31} kg	1−
Proton	1.67×10^{-27} kg	1+
Neutron	1.67×10^{-27} kg	None

*The magnitude of the charge of the electron and the proton is 1.60×10^{-19} C.

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Figure 2.21 The Periodic Table

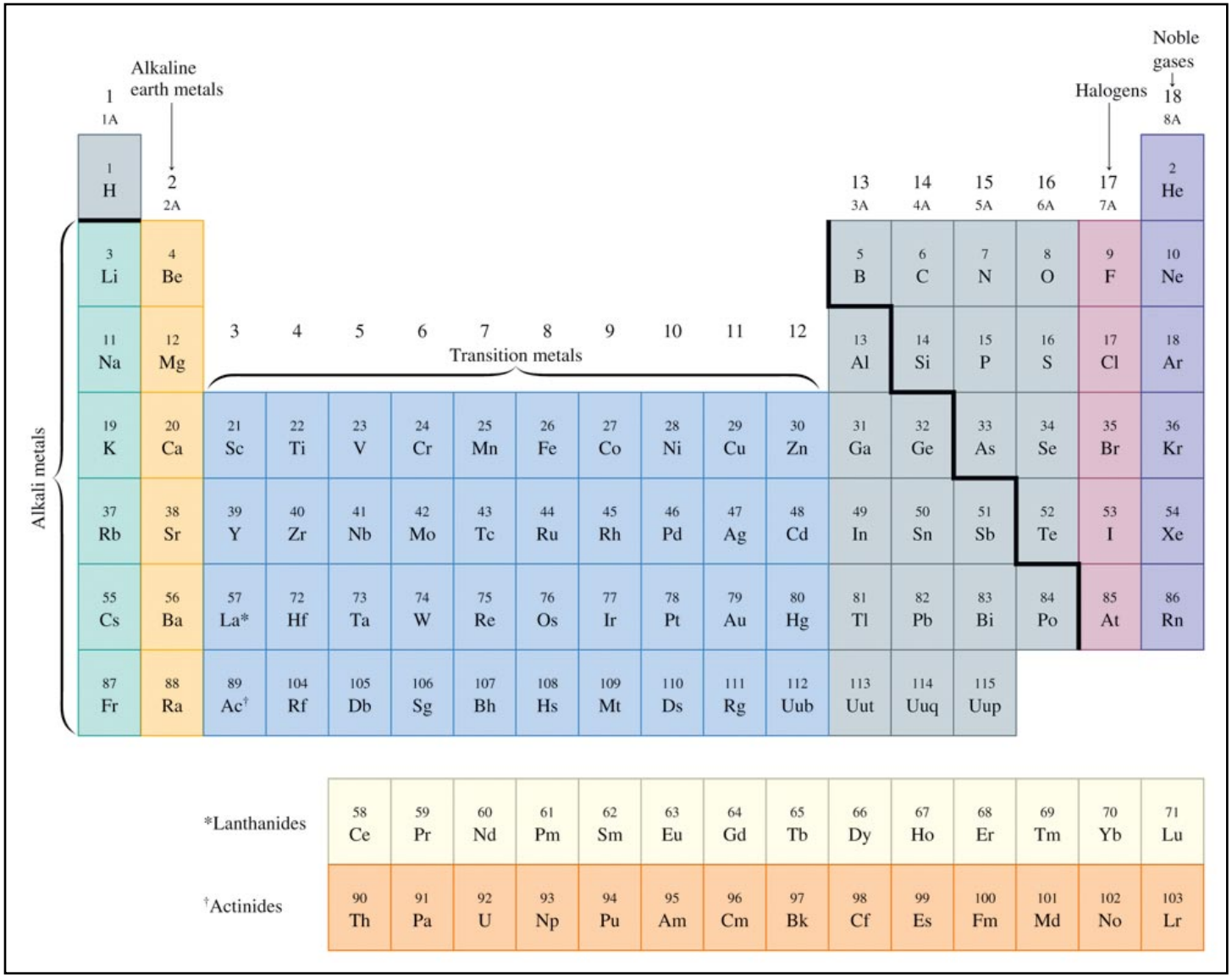

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Figure 2.23 A Flowchart for Naming Binary Compounds

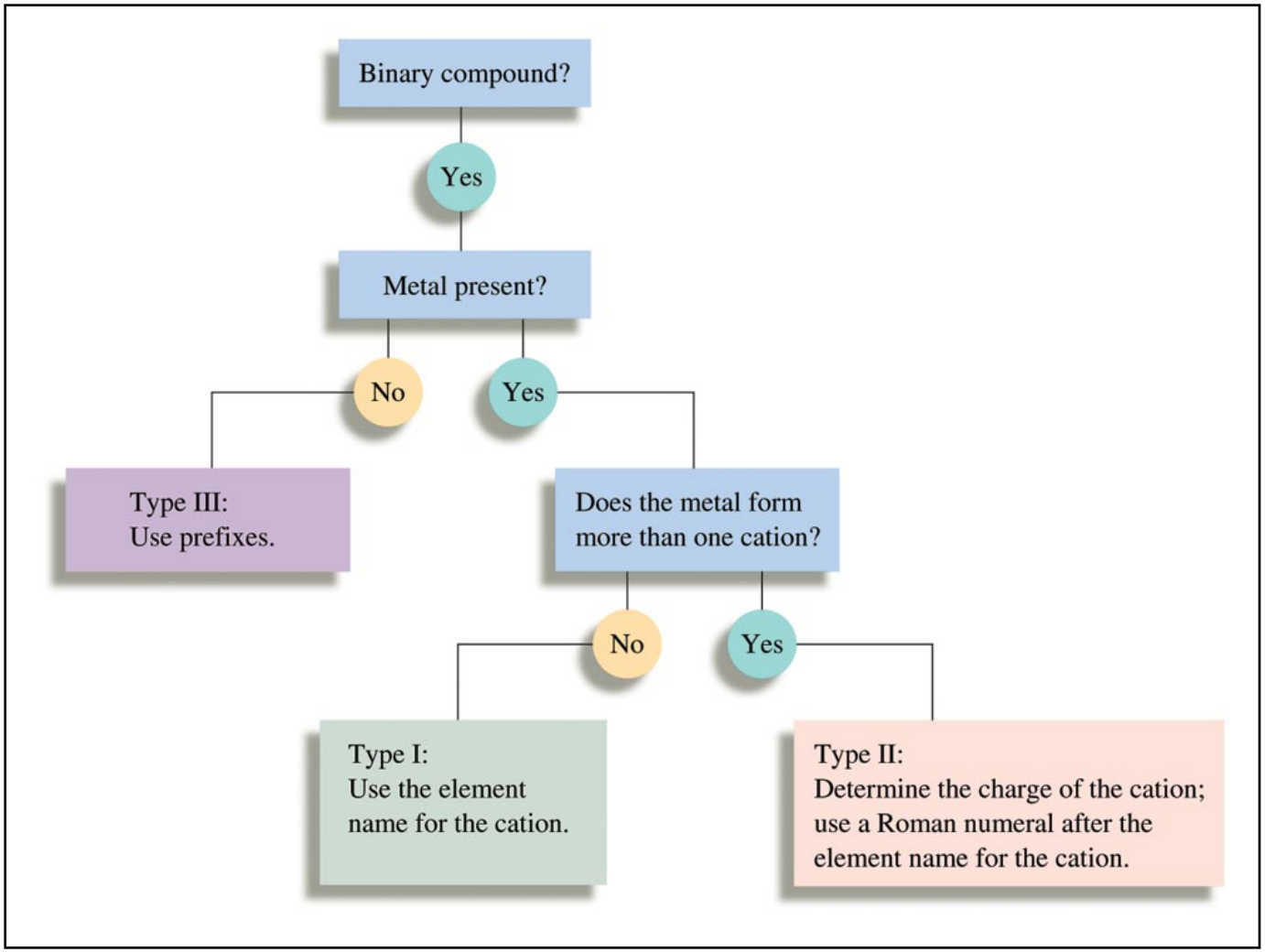

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Figure 2.24 Naming Chemical Compounds

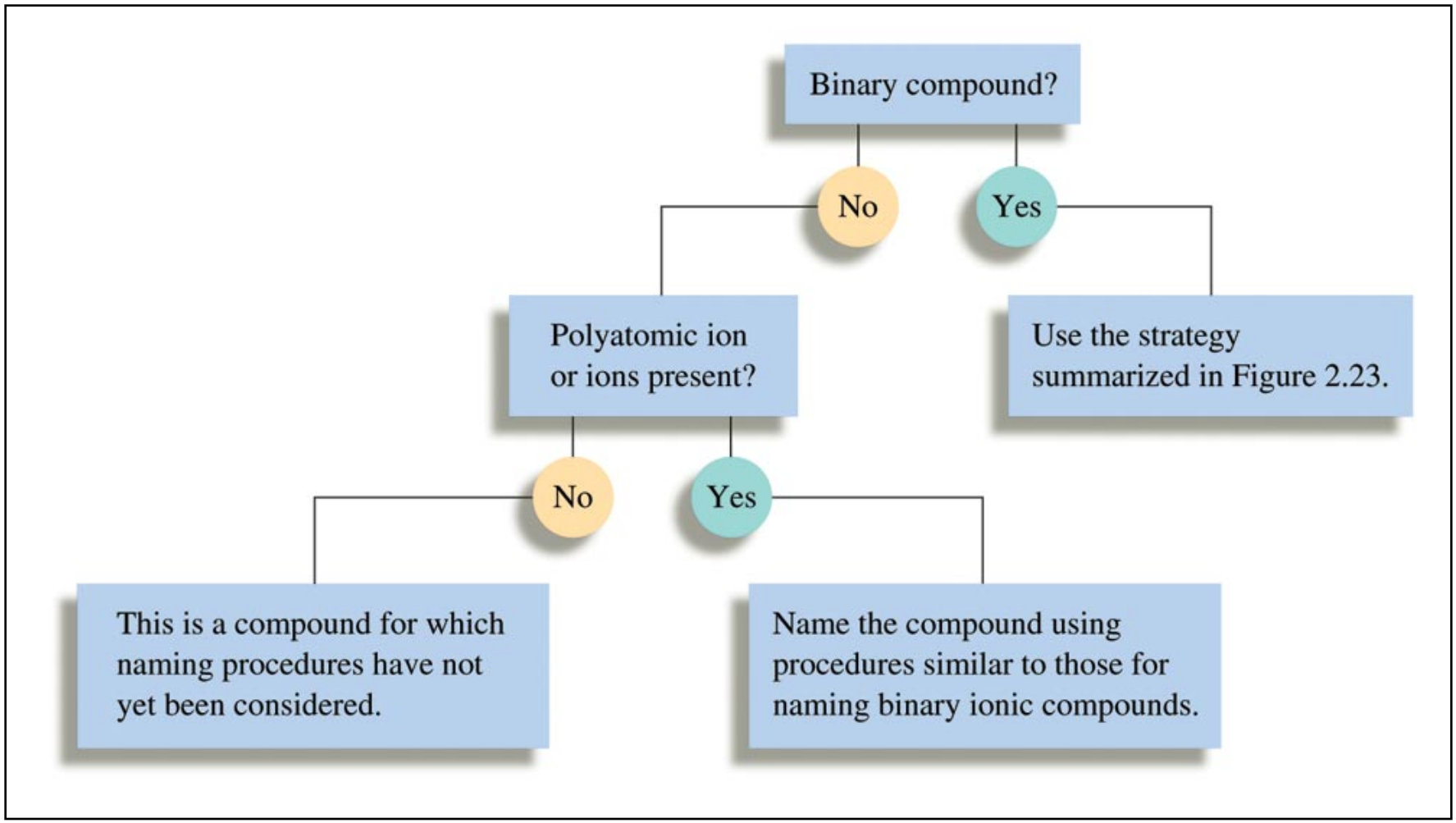

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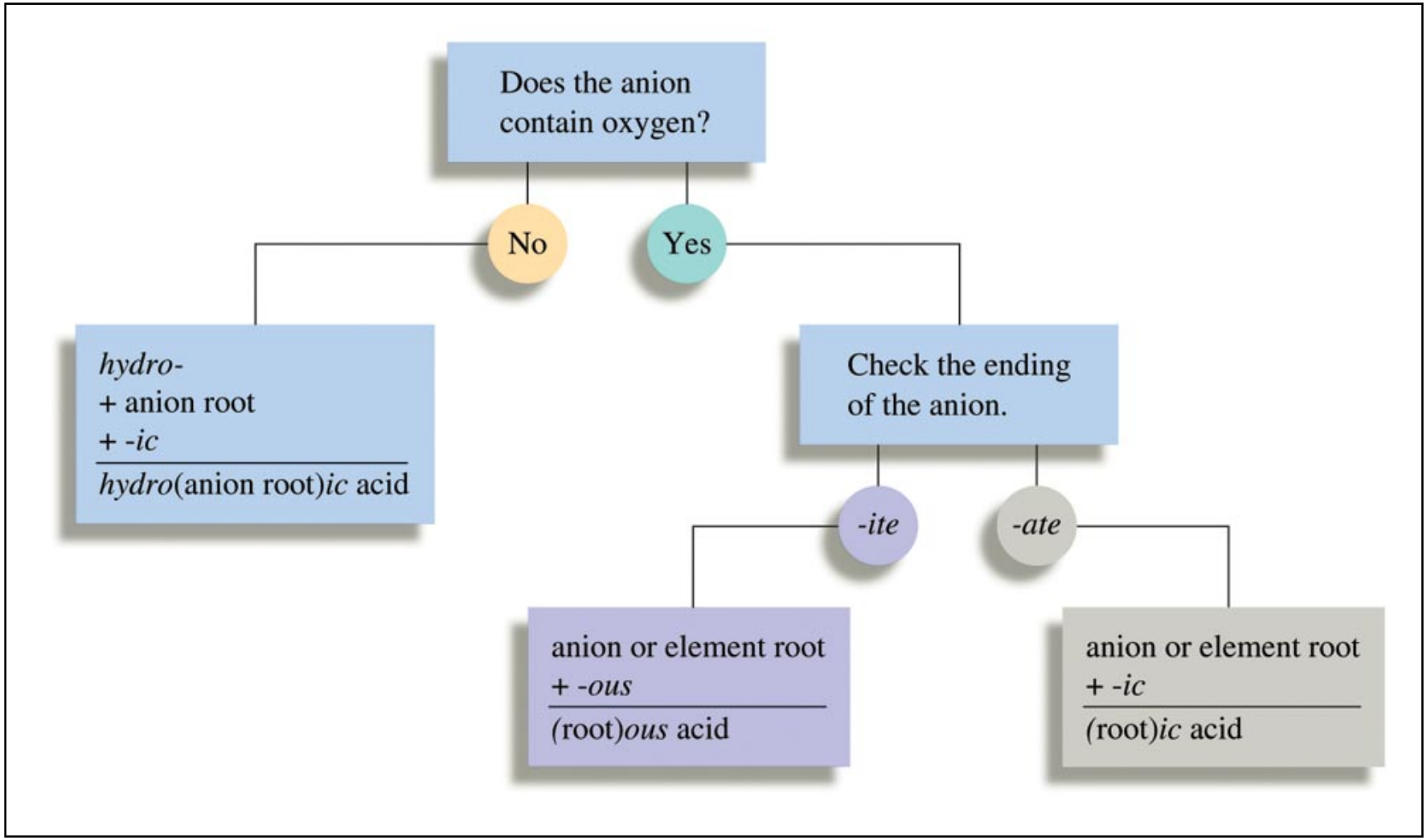
Figure 2.25 Naming Acids



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